

Please check the examination details below before entering your candidate information

Candidate surname

Other names

Centre Number

Candidate Number

Pearson Edexcel International Advanced Level

Monday 12 June 2023

Morning (Time: 1 hour 20 minutes)

Paper
reference

WCH16/01

Chemistry

International Advanced Level

UNIT 6: Practical Skills in Chemistry II

You must have:

Scientific calculator, ruler

Total Marks

Instructions

- Use **black** ink or ball-point pen.
- If pencil is used for diagrams/sketches/graphs it must be dark (HB or B).
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*

Information

- The total mark for this paper is 50.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*
- You will be assessed on your ability to organise and present information, ideas, descriptions and arguments clearly and logically, including your use of grammar, punctuation and spelling.
- A Periodic Table is printed on the back cover of this paper.

Advice

- Read each question carefully before you start to answer it.
- Show all your working in calculations and include units where appropriate.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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Answer ALL the questions. Write your answers in the spaces provided.

1 A student investigated two aqueous solutions, labelled **P** and **Q**. Both solutions were green. Each solution contained one cation and one anion.

(a) Tests were carried out on solution **P**.

Complete the table.

	Test	Observation	Inference	
(i)	A few drops of aqueous sodium hydroxide were added to 5 cm ³ of P	Chromium(III) ions may be present in P	(1)
(ii)	More sodium hydroxide solution was added to the mixture from (a)(i) until there was no further change	Chromium(III) ions are confirmed to be present in P	(1)
(iii)	A few drops of dilute nitric acid were added to 5 cm ³ of a fresh sample of P A few drops of aqueous silver nitrate were added to this acidified solution of P	A white precipitate formed	The formula of the anion likely to be responsible for the white precipitate is 	(1)

(b) State why, in the silver nitrate test on **P**, the nitric acid was not needed in this case. Justify your answer by considering the role of nitric acid in the silver nitrate test.

(2)

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(c) The student carried out tests on **Q** and inferred that it was a solution of iron(II) sulfate.

- (i) The addition of dilute aqueous ammonia to a sample of solution **Q** produced a green precipitate which changed colour on standing.

Explain why the colour change led the student to infer that **Q** contained iron(II) ions.

(2)

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- (ii) Describe a test, and its positive result, that the student could have carried out to show the presence of sulfate ions.

(2)

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- (d) Identify, by name or formula, a metal cation, other than chromium(III) and iron(II), which could give a **green** colour in an aqueous solution.

(1)

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(Total for Question 1 = 10 marks)

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2 Two organic compounds, **X** and **Y**, are colourless liquids.
Each compound contains only **one** functional group.

- (a) A few drops of deionised water are added to a beaker containing **X**.
Misty fumes are formed.

A drop of concentrated ammonia on the tip of a glass rod is placed in the misty fumes. White smoke is formed.

- (i) Deduce the functional group in **X**.
Justify your answer by referring to the observations.

(3)

- (ii) State the precaution that you would take to minimise the risk of carrying out this test on the misty fumes.
Assume gloves, safety goggles and laboratory coat are worn.

(1)

- (b) The ^{13}C NMR spectrum of **X** has two peaks.

Draw the displayed formula of **X**.

(1)

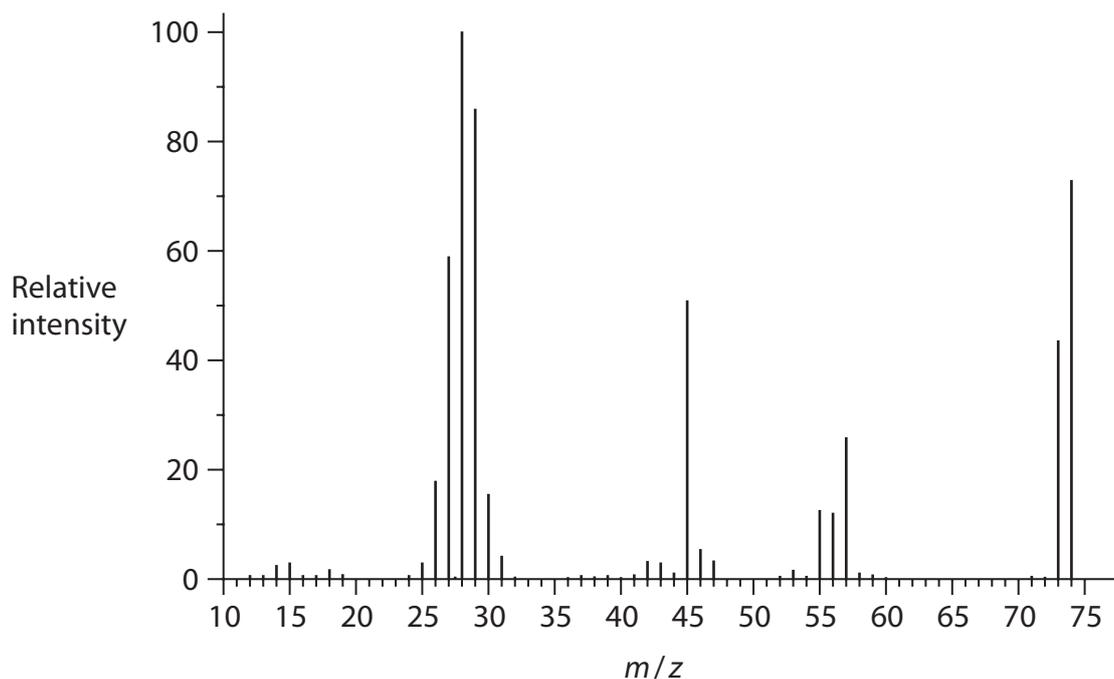


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(c) The mass spectrum of **Y** is shown.



- (i) Bubbles are observed when aqueous sodium hydrogencarbonate is added to **Y**.

Deduce the formula of the **ion** responsible for the peak at $m/z = 45$.

(1)

- (ii) Draw the structural formula of **Y**.

(1)



P 7 3 4 5 8 A 0 5 1 6

(d) Both **X** and **Y** can be used to produce esters.

(i) **Name** the compound that would react with both **X** and **Y** to form ethyl esters.

(1)

(ii) A student prepared an ester using **X** and a suitable compound.

Explain why the student added aqueous sodium hydrogencarbonate to the reaction mixture to allow the presence of an ester to be detected.

(2)

(e) Both **X** and **Y** react with concentrated ammonia but form different products.

Identify these products, by name or formula.

(2)

Product with **X**

Product with **Y**

(Total for Question 2 = 12 marks)



- 3 This question is about an experiment to investigate the kinetics of the reaction between iodine and propanone with an acid catalyst.

The equation for the reaction is shown.



To obtain the order of reaction with respect to iodine, the concentration of iodine in the reaction mixture was determined at various times.

Procedure

Step 1 Mix 25 cm³ of 1.0 mol dm⁻³ sulfuric acid with 25 cm³ of 1.0 mol dm⁻³ propanone in a beaker.

Step 2 Start a clock as 50 cm³ of 0.020 mol dm⁻³ iodine solution is added to the beaker. Mix the reactants thoroughly.

Step 3 Tip a spatula measure of sodium hydrogencarbonate into a conical flask. After 3 minutes, pipette a 10.0 cm³ sample of the reaction mixture into the conical flask and mix thoroughly.

Step 4 Titrate the iodine in the sample with 0.010 mol dm⁻³ sodium thiosulfate solution using a suitable indicator. Record the titre.

Step 5 Repeat Steps 3 and 4 every 3 minutes to obtain four more titres.

- (a) State why the sulfuric acid and propanone concentrations are both much larger than the iodine concentration.

(1)

- (b) State why sodium hydrogencarbonate is used in Step 3.

(1)

- (c) Name the indicator that would be used for the titration in Step 4, stating the colour **change** that would be seen at the end-point of the reaction.

(2)

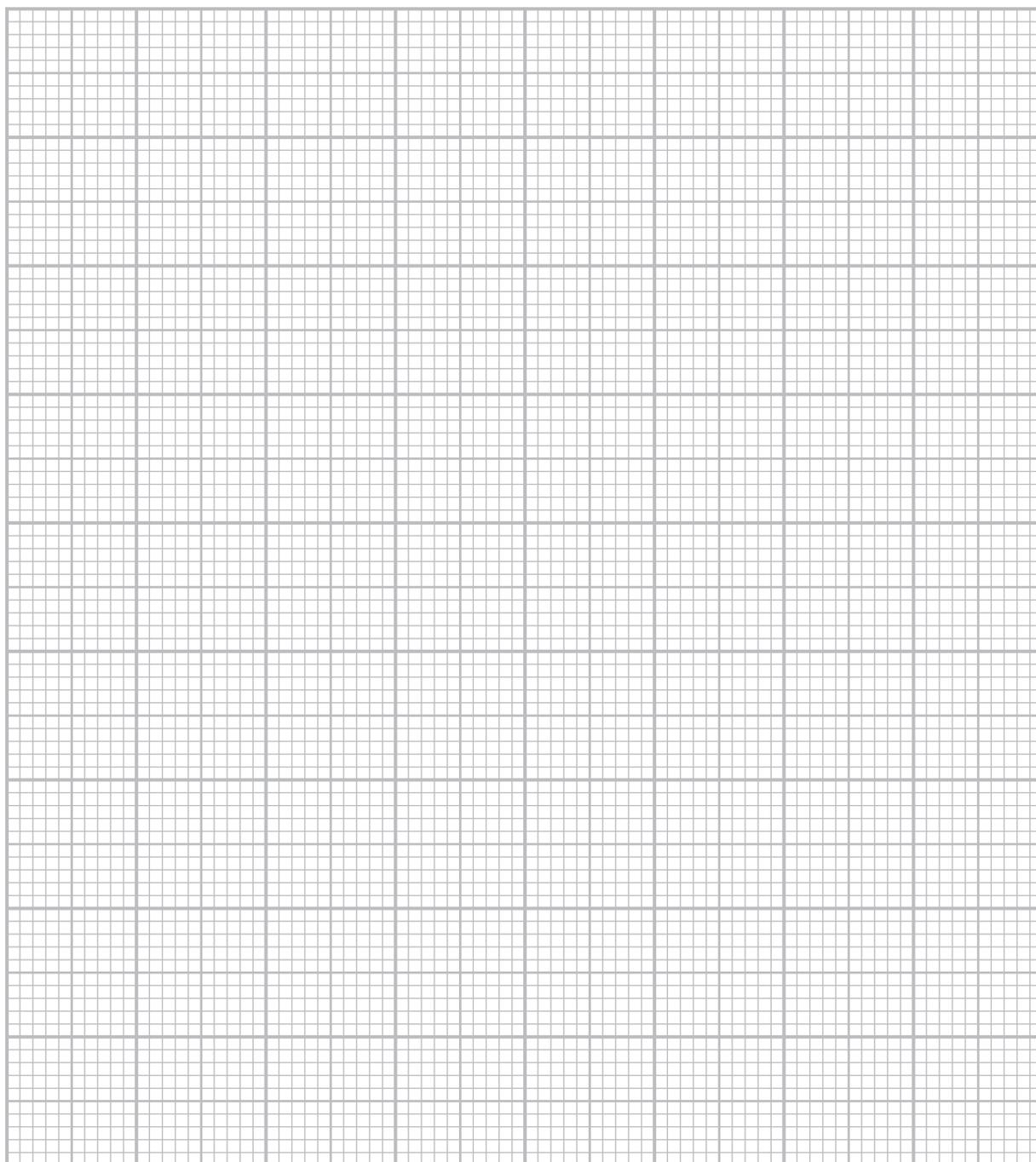


(d) Titration results from the experiment are shown.

Time/minutes	3	6	9	12	15
Titre/cm ³	16.05	15.30	14.50	13.70	12.95

(i) Plot a graph of titre against time.

(3)



(ii) State why the volume of thiosulfate may be used for plotting the graph rather than the concentration of iodine.

(1)

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(iii) State the order of reaction with respect to iodine.
Justify your answer by referring to your graph.

(1)

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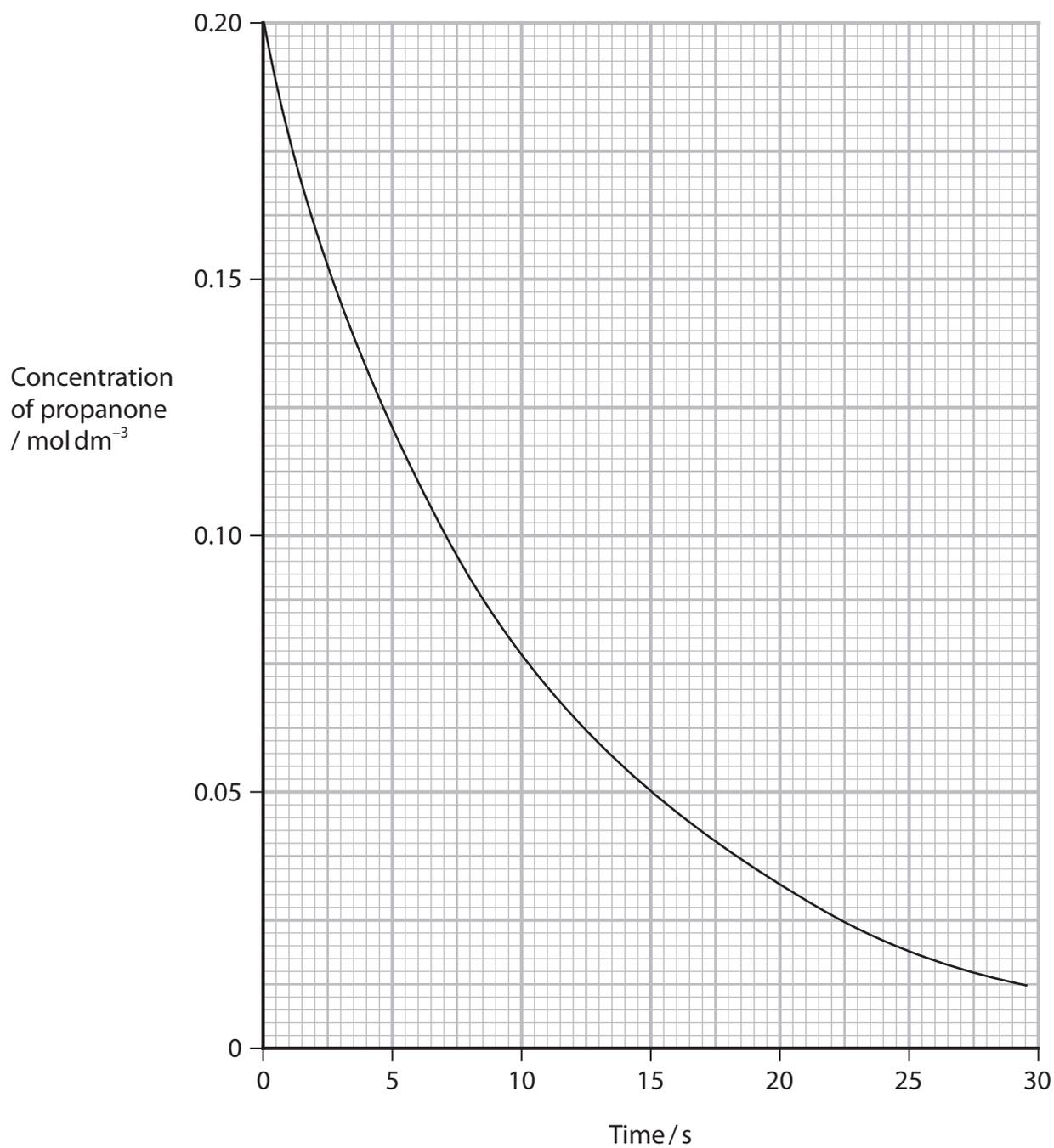
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P 7 3 4 5 8 A 0 9 1 6

(e) Further experiments were carried out to determine the reaction orders with respect to propanone and sulfuric acid.

(i) A graph of the concentration of propanone against time is shown.



The reaction is first order with respect to propanone.

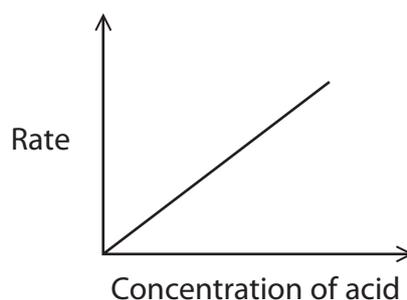
Determine two half-lives for this reaction.
You **must** show your working on the graph.

(2)

First half-life

Second half-life

- (ii) A graph of the reaction rate against the concentration of sulfuric acid is shown.



Deduce the rate equation for the overall reaction of iodine and propanone with an acid catalyst.

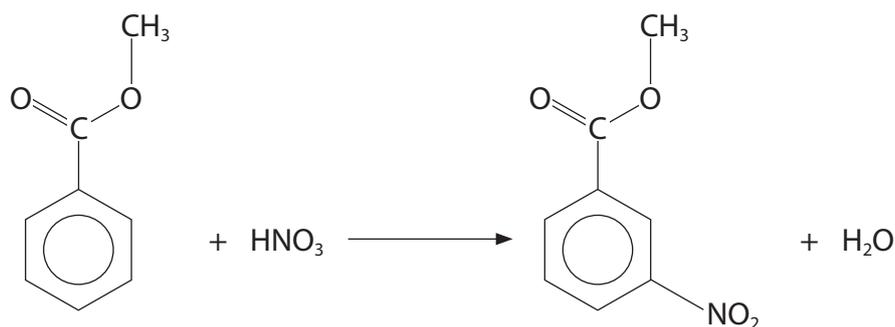
Use your answer from (d)(iii) and information from (e)(i) and the graph in (e)(ii).

(1)

(Total for Question 3 = 12 marks)



- 4 A group of students prepared methyl 3-nitrobenzoate by the nitration of methyl benzoate.



Procedure

- Step 1** Measure 9 cm^3 of concentrated sulfuric acid into a small, dry conical flask. Label the flask **A** and place it in an ice bath.
- Step 2** Add 4.0 cm^3 of methyl benzoate to flask **A**. Gently swirl the flask.
- Step 3** Mix 3 cm^3 of concentrated nitric acid and 3 cm^3 of concentrated sulfuric acid in a test tube to form the nitrating mixture. Place this test tube in the ice bath.
- Step 4** Place a thermometer in flask **A**. Add the nitrating mixture very slowly to flask **A** using a dropping pipette. Take care to ensure that the temperature of the flask contents does not rise above 15°C .
- Step 5** Remove flask **A** from the ice bath and allow it to stand at room temperature for about 10 minutes. Pour the reaction mixture into a small beaker containing crushed ice. Stir the contents of the beaker with a glass rod.
- Step 6** Allow the ice to melt. Separate the solid methyl 3-nitrobenzoate by suction filtration. Wash the solid with a small amount of deionised water and then with a little ice-cold ethanol.
- Step 7** Recrystallise the methyl 3-nitrobenzoate using ethanol as the solvent.
- Step 8** Determine the melting temperature of the purified crystals of methyl 3-nitrobenzoate.

- (a) An ice bath is a mixture of ice and water in a beaker.

Suggest an advantage of using an ice bath in Steps 1 and 3 rather than a beaker containing only ice cubes. Justify your answer.

(1)

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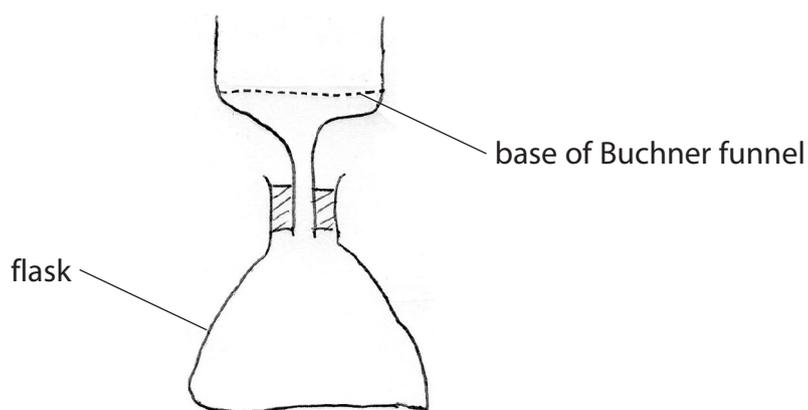


(b) Side products form if the temperature rises above 15 °C in Step 4.

Give the structure of **one** side product that may form.

(1)

(c) One student drew the suction filtration apparatus in Step 6 as shown.



Identify the **three** ways in which this diagram is incorrect.
You may assume that the apparatus is suitably clamped.

(3)

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- (d) Describe the stages in the recrystallisation of methyl 3-nitrobenzoate in Step 7, stating which stages are required to remove the insoluble and soluble impurities. Only outline details of the method are required.

(4)

- (e) The crystals must be dried before the melting temperature can be determined. Methyl 3-nitrobenzoate cannot be dried by the addition of a solid drying agent such as anhydrous calcium chloride.

- (i) Suggest why the addition of a solid drying agent is not suitable to dry methyl 3-nitrobenzoate.

(1)

- (ii) State how the crystals of methyl 3-nitrobenzoate could be dried.

(1)



(f) The mass of dry methyl 3-nitrobenzoate crystals prepared by one of the students was 3.05 g.

Calculate the percentage yield by mass of methyl 3-nitrobenzoate using the data shown.

Compound	Molar mass / g mol ⁻¹	Density / g cm ⁻³
methyl benzoate	136	1.08
methyl 3-nitrobenzoate	181	

(3)

(g) The melting temperature range of methyl 3-nitrobenzoate is given in a data book as 78–80 °C.

Suggest a melting temperature **range** for a sample of the methyl 3-nitrobenzoate **before** recrystallisation. Justify your answer.

(2)

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(Total for Question 4 = 16 marks)

TOTAL FOR PAPER = 50 MARKS

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P 7 3 4 5 8 A 0 1 5 1 6

The Periodic Table of Elements

1 2 3 4 5 6 7 0 (8) (18)

1.0	H	hydrogen	1
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Key

relative atomic mass
atomic symbol
name
atomic (proton) number

(1)	(2)	(3)	(4)	(5)	(6)	(7)	(8)	(9)	(10)	(11)	(12)	(13)	(14)	(15)	(16)	(17)	(18)
6.9	9.0	45.0	47.9	50.9	52.0	54.9	55.8	58.9	58.7	63.5	65.4	10.8	12.0	14.0	16.0	19.0	4.0
Li	Be	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	B	C	N	O	F	He
lithium	beryllium	scandium	titanium	vanadium	chromium	manganese	iron	cobalt	nickel	copper	zinc	boron	carbon	nitrogen	oxygen	fluorine	helium
3	4	21	22	23	24	25	26	27	28	29	30	5	6	7	8	9	2
23.0	24.3	88.9	91.2	92.9	95.9	[98]	101.1	102.9	106.4	107.9	112.4	27.0	28.1	31.0	32.1	35.5	39.9
Na	Mg	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	Al	Si	P	S	Cl	Ar
sodium	magnesium	yttrium	zirconium	niobium	molybdenum	technetium	ruthenium	rhodium	palladium	silver	cadmium	aluminium	silicon	phosphorus	sulfur	chlorine	argon
11	12	39	40	41	42	43	44	45	46	47	48	13	14	15	16	17	18
39.1	40.1	88.9	91.2	92.9	95.9	[98]	101.1	102.9	106.4	107.9	112.4	69.7	72.6	74.9	79.0	79.9	83.8
K	Ca	La*	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Ga	Ge	As	Se	Br	Kr
potassium	calcium	lanthanum	hafnium	tantalum	tungsten	rhenium	osmium	iridium	platinum	gold	mercury	gallium	germanium	arsenic	selenium	bromine	krypton
19	20	57	72	73	74	75	76	77	78	79	80	31	32	33	34	35	36
85.5	87.6	138.9	178.5	180.9	183.8	186.2	190.2	192.2	195.1	197.0	200.6	69.7	72.6	74.9	79.0	79.9	131.3
Rb	Sr	La*	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	In	Sn	Sb	Te	I	Xe
rubidium	strontium	lanthanum	hafnium	tantalum	tungsten	rhenium	osmium	iridium	platinum	gold	mercury	indium	tin	antimony	tellurium	iodine	xenon
37	38	57	72	73	74	75	76	77	78	79	80	49	50	51	52	53	54
132.9	137.3	138.9	178.5	180.9	183.8	186.2	190.2	192.2	195.1	197.0	200.6	114.8	118.7	121.8	127.6	126.9	222
Cs	Ba	La*	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Pb	Bi	Po	At	Rn	Rn
caesium	barium	lanthanum	hafnium	tantalum	tungsten	rhenium	osmium	iridium	platinum	gold	mercury	lead	bismuth	polonium	astatine	radon	radon
55	56	57	72	73	74	75	76	77	78	79	80	82	83	84	85	86	86
[223]	[226]	[227]	[261]	[262]	[266]	[264]	[277]	[268]	[271]	[272]	[272]	204.4	207.2	209.0	[210]	[222]	[222]
Fr	Ra	Ac*	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Rg	Tl	Pb	Bi	Po	At	Rn
francium	radium	actinium	rutherfordium	dubnium	seaborgium	bohrium	hassium	meitnerium	damascium	roentgenium	roentgenium	thallium	lead	bismuth	polonium	astatine	radon
87	88	89	104	105	106	107	108	109	110	111	111	81	82	83	84	85	86

Elements with atomic numbers 112-116 have been reported but not fully authenticated

140	141	144	150	152	157	163	165	167	169	173	175
Ce	Pr	Nd	Sm	Eu	Gd	Dy	Ho	Er	Tm	Yb	Lu
cerium	praseodymium	neodymium	samarium	europium	gadolinium	dysprosium	holmium	erbium	thulium	ytterbium	lutetium
58	59	60	62	63	64	66	67	68	69	70	71
232	[231]	238	[242]	[243]	[247]	[251]	[254]	[253]	[256]	[254]	[257]
Th	Pa	U	Pu	Am	Cm	Cf	Es	Fm	Md	No	Lr
thorium	protactinium	uranium	plutonium	americium	curium	californium	einsteinium	fermium	mendeleevium	nobelium	lawrencium
90	91	92	94	95	96	98	99	100	101	102	103

* Lanthanide series
* Actinide series

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